

## Chemistry 338 2<sup>nd</sup> Semester Final Exam Review Sheet

### District Assessment

Element	Symbol	Atomic #	Protons	Neutrons	Electrons	Mass #	Charge
Tungsten	<sup>184</sup> W <sub>74</sub>	74	74	110	72	184	+2

1. Name the compound
  - a. Fe<sub>2</sub>O<sub>3</sub> – *iron (III) oxide*
  - b. Cl<sub>2</sub>O<sub>8</sub> – *dichlorine octoxide*
  
2. Write the formula
  - a. Trisulfur Dinitride – *S<sub>3</sub>N<sub>2</sub>*
  - b. Zinc Hydroxide – *Zn(OH)<sub>2</sub>*
  - c. Tin (IV) oxide – *SnO<sub>2</sub>*
  
3. The pressure of a gas is 843 mm Hg at 24°C. What will the pressure be at standard temperature? *Volume is constant.*

$$\frac{P_1}{T_1} = \frac{P_2}{T_2} \quad \frac{843 \text{ mmHg}}{297\text{K}} = \frac{P_2}{273\text{K}} \quad P_2 = 775 \text{ mm Hg}$$
  
4. The volume of a gas is 81.7 mL when the pressure is 119 kPa and the temperature is 73°C. What is the volume at STP?
 
$$\frac{P_1V_1}{T_1} = \frac{P_2V_2}{T_2} \quad \frac{(119\text{kPa})(81.7\text{mL})}{346\text{K}} = \frac{(101.3\text{kPa})(V_2)}{273\text{K}} \quad V_2 = 75.7 \text{ mL}$$
  
5. Between Be and Ba, which has a larger ionization energy? Larger atomic radius?
 

*Larger ionization energy – Be*  
*Larger atomic radius – Ba*
  
6. Balance the equation & classify the type of reaction:
 
$$2\text{Al} + 3 \text{CuSO}_4 \rightarrow 1\text{Al}_2(\text{SO}_4)_3 + 3 \text{Cu} \quad \text{single replacement}$$
  
7. Predict the products & classify the type of reaction:
 
$$\text{H}_3\text{PO}_4 + 3 \text{NaOH} \rightarrow 3 \text{H}_2\text{O} + \text{Na}_3\text{PO}_4 \quad \text{double replacement}$$
  
8. Calculate the Molar Mass: (NH<sub>4</sub>)<sub>2</sub>CO<sub>3</sub>  
*96 g/mole*
  
9. Convert 2.75 moles of N<sub>2</sub>O<sub>5</sub> to molecules
 
$$\frac{2.75 \text{ moles}}{1} \times \frac{6.02 \times 10^{23} \text{ molecules}}{1 \text{ mole}} = 1.66 \times 10^{24} \text{ molecules N}_2\text{O}_5$$

### Solutions

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