

Stoichiometry worksheet #2

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1. The equation for the synthesis of ammonia is below. How many moles of H₂ are needed to produce 9.33 moles of NH₃?



$$\frac{9.33 \text{ mol NH}_3}{1} \times \frac{3 \text{ mol H}_2}{2 \text{ mol NH}_3} = \boxed{13.995 \text{ mol H}_2}$$

2. What mass (in grams) of copper is produced when 3.80 g of iron reacts with excess CuSO₄?



$$\frac{3.80 \text{ g Fe}}{1} \times \frac{1 \text{ mol Fe}}{55.845 \text{ g Fe}} \times \frac{1 \text{ mol Cu}}{1 \text{ mol Fe}} \times \frac{63.546 \text{ g Cu}}{63.546 \text{ g Cu}} = \boxed{4.32 \text{ g Cu}}$$

3. If 5.00g of CH₄ are burned according to the equation below, what mass in grams of CO₂ are produced?



$$\frac{5.00 \text{ g CH}_4}{1} \times \frac{1 \text{ mol CH}_4}{16.043 \text{ g CH}_4} \times \frac{1 \text{ mol CO}_2}{1 \text{ mol CH}_4} \times \frac{44.01 \text{ g CO}_2}{44.01 \text{ g CO}_2} = \boxed{13.7 \text{ g CO}_2}$$

4. In the reaction that follows, determine how many grams of aluminum chloride (AlCl₃) will be produced from 92.01 g of Cl₂.



$$\frac{92.01 \text{ g Cl}_2}{1} \times \frac{1 \text{ mol Cl}_2}{70.906 \text{ g Cl}_2} \times \frac{2 \text{ mol AlCl}_3}{3 \text{ mol Cl}_2} \times \frac{133.34 \text{ g AlCl}_3}{1 \text{ mol AlCl}_3} = \boxed{115.4 \text{ g AlCl}_3}$$

Chemistry Stoichiometry Problem Sheet 2 Answers Key

Jicheng Xie



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