

**Study Tip**

**Organize New Information** Different people learn in different ways. If you learn best visually, draw concept maps, outlines, and diagrams to help you organize the chapter content. If you learn best by listening, take turns reading the chapter aloud with a partner.



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**Key Concepts****10.1 The Mole: A Measurement of Matter**

- Three methods for measuring the amount of a substance are by count, by mass, and by volume.
- A mole of any substance always contains Avogadro's number of representative particles, or  $6.02 \times 10^{23}$  representative particles.
- The atomic mass of an element expressed in grams is the mass of a mole of the element.
- To calculate the molar mass of a compound, find the number of grams of each element contained in one mole of the compound. Then add the masses of the elements in the compound.

**10.2 Mole–Mass and Mole–Volume Relationships**

- The molar mass of an element or compound is the conversion factor for converting between the mass and the number of moles of a substance.

- One mole of any gas occupies a volume of 22.4 L at standard temperature and pressure. One mole of any substance contains Avogadro's number of particles, so 22.4 L of any gas at STP contains  $6.02 \times 10^{23}$  representative particles of that gas.

**10.3 Percent Composition and Chemical Formulas**

- To determine the percent by mass of any element in a given compound, divide the element's mass by the mass of the compound and multiply by 100%.
- An empirical formula of a compound is the simplest whole-number ratio of atoms of the elements in the compound.
- The molecular formula of a compound is either the same as its experimentally determined empirical formula, or it is a simple whole-number multiple of it.

**Vocabulary**

- Avogadro's hypothesis (p. 300)
- Avogadro's number (p. 290)
- empirical formula (p. 300)
- molar mass (p. 294)
- molar volume (p. 300)
- mole (mol) (p. 290)
- percent composition (p. 300)
- representative particle (p. 290)
- standard temperature and pressure (STP) (p. 300)

**Key Equations**

$$\bullet \text{ moles} = \text{representative particles} \times \frac{1 \text{ mole}}{6.02 \times 10^{23} \text{ representative particles}}$$

$$\bullet \text{ representative particles} = \text{moles} \times \frac{6.02 \times 10^{23} \text{ representative particles}}{1 \text{ mole}}$$

$$\bullet \text{ mass (grams)} = \text{number of moles} \times \frac{\text{mass (grams)}}{1 \text{ mole}}$$

$$\bullet \text{ moles} = \text{mass (grams)} \times \frac{1 \text{ mole}}{\text{mass (grams)}}$$

$$\bullet \frac{\text{grams}}{\text{mole}} = \frac{\text{grams}}{1} \times \frac{22.4 \text{ L}}{1 \text{ mole}}$$

$$\bullet \text{ volume of gas} = \text{moles of gas} \times \frac{22.4 \text{ L}}{1 \text{ mole}}$$

$$\bullet \% \text{ mass of element} = \frac{\text{mass of element}}{\text{mass of compound}} \times 100\%$$

**Organizing Information**

Organize the major ideas of the chapter.



**Concept Map 10:** Solve the Concept Map with the help of an interactive tutorial.

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314 Chapter 10

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# Chemistry Chapter 10 Study Guide

**DJ Losen**



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## **Table of Contents Chemistry Chapter 10 Study Guide**

1. Understanding the eBook Chemistry Chapter 10 Study Guide
  - The Rise of Digital Reading Chemistry Chapter 10 Study Guide
  - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Chapter 10 Study Guide
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chemistry Chapter 10 Study Guide
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Chapter 10 Study Guide
  - Personalized Recommendations
  - Chemistry Chapter 10 Study Guide User Reviews and Ratings
  - Chemistry Chapter 10 Study Guide and Bestseller Lists

5. Accessing Chemistry Chapter 10 Study Guide Free and Paid eBooks
  - Chemistry Chapter 10 Study Guide Public Domain eBooks
  - Chemistry Chapter 10 Study Guide eBook Subscription Services
  - Chemistry Chapter 10 Study Guide Budget-Friendly Options
6. Navigating Chemistry Chapter 10 Study Guide eBook Formats
  - ePub, PDF, MOBI, and More
  - Chemistry Chapter 10 Study Guide Compatibility with Devices
  - Chemistry Chapter 10 Study Guide Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chemistry Chapter 10 Study Guide
  - Highlighting and Note-Taking Chemistry Chapter 10 Study Guide
  - Interactive Elements Chemistry Chapter 10 Study Guide
8. Staying Engaged with Chemistry Chapter 10 Study Guide
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chemistry Chapter 10 Study Guide
9. Balancing eBooks and Physical Books Chemistry Chapter 10 Study Guide
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chemistry Chapter 10 Study Guide
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Chemistry Chapter 10 Study Guide
  - Setting Reading Goals Chemistry Chapter 10 Study Guide
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemistry Chapter 10 Study Guide
  - Fact-Checking eBook Content of Chemistry Chapter 10 Study Guide
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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