



# CHEMISTRY CHEAT SHEET



# EQUILIBRIUM

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Chemical Equilibrium occurs in a reversible reaction when the rate of the forward reaction equals the rate of the reverse reaction. Equilibrium constants are represented by the symbol  $K_c$ .

**What is Equilibrium?**  
 A reversible reaction is one that can proceed in either direction. At equilibrium, the rate of the forward reaction is equal to the rate of the reverse reaction. The concentrations of the reactants and products remain constant over time.

**Equilibrium Constant  $K_c$**   
 The equilibrium constant  $K_c$  is a ratio of the concentrations of the products to the concentrations of the reactants, each raised to the power of their stoichiometric coefficients. It is a constant for a given reaction at a given temperature.



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Reaction	Equilibrium Constant $K_c$
$aA + bB \rightleftharpoons cC + dD$	$K_c = \frac{[C]^c [D]^d}{[A]^a [B]^b}$
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Factor	Effect on Equilibrium	Example
Concentration	Increasing the concentration of a reactant shifts the equilibrium to the right, favoring the products. Increasing the concentration of a product shifts the equilibrium to the left, favoring the reactants.	$N_2 + 3H_2 \rightleftharpoons 2NH_3$ Adding $N_2$ shifts the equilibrium to the right, producing more $NH_3$ .
Pressure	Increasing the pressure shifts the equilibrium to the side with fewer moles of gas. Decreasing the pressure shifts the equilibrium to the side with more moles of gas.	$N_2 + 3H_2 \rightleftharpoons 2NH_3$ Increasing the pressure shifts the equilibrium to the right, producing more $NH_3$ .
Temperature	Increasing the temperature shifts the equilibrium to the endothermic side. Decreasing the temperature shifts the equilibrium to the exothermic side.	$N_2 + 3H_2 \rightleftharpoons 2NH_3$ The forward reaction is exothermic. Increasing the temperature shifts the equilibrium to the left, favoring the reactants.

# Chemistry Equilibrium Study Guide

**Chao Zhang**



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Jul 6 2021 this introductory lesson introduces the key kingdoms of study discusses toby green s misconceptions of african history and includes a map activity to allow students to