

Acid Base Practice Questions

- 1) Write equations for the dissociation of the following compounds in aqueous solution. Identify the acid base conjugate pairs in each case (use appropriate arrows for strong and weak species)
- $\text{CH}_3\text{COOH} + \text{H}_2\text{O} \leftrightarrow \text{CH}_3\text{COO}^- + \text{H}_3\text{O}^+$
 - $\text{CH}_3\text{NH}_2 + \text{H}_2\text{O} \leftrightarrow \text{CH}_3\text{NH}_3^+ + \text{OH}^-$
 - $\text{NH}_3 + \text{H}_2\text{O} \leftrightarrow \text{NH}_4^+ + \text{OH}^-$
 - $\text{HNO}_3 + \text{H}_2\text{O} \leftrightarrow \text{NO}_3^- + \text{H}_3\text{O}^+$
- 2) Define the term amphiprotic. Write appropriate equations to show the amphiprotic nature of water.
Can accept or donate a proton
- $\text{NH}_3 + \text{H}_2\text{O} \leftrightarrow \text{NH}_4^+ + \text{OH}^-$ (acting as an acid)
 - $\text{HNO}_3 + \text{H}_2\text{O} \leftrightarrow \text{NO}_3^- + \text{H}_3\text{O}^+$ (acting as a base)
- 3) Calculate the pH of the following solutions
- $[\text{H}^+] = 2.5 \times 10^{-5} \text{ mol/L}$ 4.6
 - $[\text{OH}^-] = 1.8 \times 10^{-4} \text{ mol/L}$ 10.25
 - 0.006 M NaOH 11.77
- 4) Compare the $[\text{H}^+]$ in solutions with the pH of 3 and 5 (pH 3 has 100x higher $[\text{H}^+]$)
- 5) Calculate the concentration of H^+ and OH^- ions in 0.005 M solution of NaOH.
- $[\text{OH}^-] = 0.005$
 - $[\text{H}^+] = 2.0 \times 10^{-12}$
- 6) Calculate the pH of a 0.0004 M solution of $\text{Ba}(\text{OH})_2$ 10.9
- 7) Describe how you would experimentally distinguish a strong base and a strong acid.
- Test pH
 - React with a metal
- 8) Consider the reaction: $2\text{NH}_3 + \text{H}_2\text{SO}_4 \rightarrow (\text{NH}_4)_2\text{SO}_4$
- What volume of 2.0 M ammonia is needed to neutralize 50 mL of a 0.5 M sulfuric acid? 25mL
 - Is the ammonium sulphate an acidic, basic or a neutral salt? Acidic
- 9) What are indicators? What factors must be considered when selecting an indicator for a titration?
- Substances that change colour over a specific pH range (called endpoint of indicator)
 - The pH at equivalence for titration must fall within the endpoint of indicator
- 10) Calculate the pH of (pH of a weak acid)
- 0.005 M solution of ethanoic acid ($K_a = 1.76 \times 10^{-5}$) 3.52
 - 0.003 M solution of ammonia ($K_b = 1.8 \times 10^{-5}$) 10.4
- 11) Calculate the
- pKa for acetic acid ($K_a = 1.76 \times 10^{-5}$) 4.75
 - pKb for NH_3 ($K_b = 1.8 \times 10^{-5}$) 4.74
- 12) How are K_a , pKa and K_b and pKb related? (higher K_a , lower pKa) (higher K_b , lower pKb)
- 13) What are buffers?
- Substances that resist change in pH when small amounts of a strong base/acid is added to them
- 14) What are the main components of an acidic buffer, a basic buffer? Give examples.
- Acidic buffer: weak acid and a salt containing its conjugate base
 - Basic Buffer: weak base and a salt containing its conjugate acid

Acid Base Equilibrium Practice Problems Multiple Choice

David Baud



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