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CHAPTER

15

STUDY GUIDE FOR CONTENT MASTERY

Solutions

Section 15.1 What are solutions?

In your textbook, read about the characteristics of solutions.

Use each of the terms below just once to complete the passage.

immiscible	liquid	soluble	solution
insoluble	miscible	solvent	solvent

Air is a(n) **(1)** solution of oxygen gas dissolved in nitrogen gas. The oxygen is an **(2)** solute, and nitrogen is the **(3)** solvent. Because oxygen gas dissolves in a solvent, oxygen gas is an **(4)** soluble substance. A substance that does not dissolve is **(5)** insoluble. **(6)** Liquid solutions are the most common type of solutions. If one liquid is soluble in another liquid, such as acetic acid in water, the two liquids are **(7)** miscible. However, if one liquid is insoluble in another, the liquids are **(8)** immiscible.

Read about hydration in aqueous solutions in your textbook.

The diagram shows the hydration of solid sodium chloride to form an aqueous solution. Use the diagram to answer the following questions.



9. Hydration is hydration in which the solvent is water. What is hydration?
Solvation is the process of surrounding solute particles with solvent particles to form a solution.

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85

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CHAPTER

15

Section 15.1 continued

16. An sodium chloride dissolves in water. What happens to the sodium and chloride ions? **The sodium and chloride ions are separated and surrounded by the water molecules.**

17. Explain the orientation of the water molecules around the sodium ions and chloride ions. **Because the sodium ion is positively charged, it attracts the negatively charged portion of the water molecule (the oxygen atom) and repels the positively charged portion of the water molecule (the hydrogen atoms). Because the chloride ion is negatively charged, it attracts the positively charged portion of the water molecule and repels the negatively charged portion.**

18. How does the strength of the attraction between water molecules and sodium and chloride ions compare with the strength of the attraction between the sodium ions and chloride ions? **How do you know?**

The attraction between the water molecules and the sodium and chloride ions is greater than the attraction between the sodium and chloride ions. The greater strength of attraction between the water molecules and the ions is what causes the solvation process to occur.

19. List three ways that the rate of solvation may be increased.

stirring or shaking the solution, breaking the solute into smaller pieces, and heating the solvent

In your textbook, read about heat of solution, solubility, and factors that affect solubility.

For each statement below, write *true* or *false*.

- true** 14. The overall energy change that occurs when a solution forms is called the heat of solution.
- false** 15. Solubility is a measure of the maximum amount of solute that dissolves in a given amount of solvent at a specified temperature and pressure.
- false** 16. Solvation continues as long as the solvation rate is less than the crystallization rate.
- true** 17. In a saturated solution, solvation and crystallization are in equilibrium.
- true** 18. Additional water can be dissolved in an unsaturated solution.
- true** 19. The solubility of a gas dissolved in a liquid decreases as the temperature of the solution increases.

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86

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Didier Musso



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Table of Contents Chapter Study Guide For Content Mastery Answer Key Chemistry

1. Understanding the eBook Chapter Study Guide For Content Mastery Answer Key Chemistry
 - The Rise of Digital Reading Chapter Study Guide For Content Mastery Answer Key Chemistry
 - Advantages of eBooks Over Traditional Books
2. Identifying Chapter Study Guide For Content Mastery Answer Key Chemistry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chapter Study Guide For Content Mastery Answer Key Chemistry
 - User-Friendly Interface

4. Exploring eBook Recommendations from Chapter Study Guide For Content Mastery Answer Key Chemistry
 - Personalized Recommendations
 - Chapter Study Guide For Content Mastery Answer Key Chemistry User Reviews and Ratings
 - Chapter Study Guide For Content Mastery Answer Key Chemistry and Bestseller Lists
5. Accessing Chapter Study Guide For Content Mastery Answer Key Chemistry Free and Paid eBooks
 - Chapter Study Guide For Content Mastery Answer Key Chemistry Public Domain eBooks
 - Chapter Study Guide For Content Mastery Answer Key Chemistry eBook Subscription Services
 - Chapter Study Guide For Content Mastery Answer Key Chemistry Budget-Friendly Options
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 - ePub, PDF, MOBI, and More
 - Chapter Study Guide For Content Mastery Answer Key Chemistry Compatibility with Devices
 - Chapter Study Guide For Content Mastery Answer Key Chemistry Enhanced eBook Features
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 - Highlighting and Note-Taking Chapter Study Guide For Content Mastery Answer Key Chemistry
 - Interactive Elements Chapter Study Guide For Content Mastery Answer Key Chemistry
8. Staying Engaged with Chapter Study Guide For Content Mastery Answer Key Chemistry
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chapter Study Guide For Content Mastery Answer Key Chemistry
9. Balancing eBooks and Physical Books Chapter Study Guide For Content Mastery Answer Key Chemistry
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter Study Guide For Content Mastery Answer Key Chemistry
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chapter Study Guide For Content Mastery Answer Key Chemistry
 - Setting Reading Goals Chapter Study Guide For Content Mastery Answer Key Chemistry
 - Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Chapter Study Guide For Content Mastery Answer Key Chemistry
 - Fact-Checking eBook Content of Chapter Study Guide For Content Mastery Answer Key Chemistry
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
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