

**CHM\_160**

**Worksheet\_Chapter\_12\_answers**

**Unit cells and basic structure, crystalline solids, ionic solids**

$4r = 2l, l\sqrt{2}$  or  $l\sqrt{3}$  ;  $V = (4/3)\pi r^3$  ;

**packing fraction = 0.524, 0.680, 0.740** ,  $\rho = (\# \text{ atoms/U.C})(M.M)/N_A$

**Answer the following questions:**

1. Lead has an atomic radius of 175 pm and a face-centered crystal structure. The molar mass of lead is 207.2 g/mol. Calculate the density of lead in g/cm<sup>3</sup>.

**Answer: 11.3 g/cm<sup>3</sup>**

2. Iron, which has a molar mass of 55.847 g/mol, crystallizes in a body-centered cubic unit cell. The atomic radius of iron is 124.1 pm. Use this information to calculate the density of iron in g/cm<sup>3</sup>.

**Answer: 7.879 g/cm<sup>3</sup>**

3. An element crystallizes in the face-centered cubic unit cell. The **volume of an atom** of the element is  $1.05 \times 10^7$  pm<sup>3</sup>, and the density of the element is 22.55 g/cm<sup>3</sup>. Calculate the molar mass of the element.

**Answer: 193 g/mol**

4. An element crystallizes with a body-centered cubic unit cell. The atomic radius of the element is 136 pm. The density of the element is 10.3 g/cm<sup>3</sup>. Calculate the mass of a single atom of the element in grams.

**Answer :  $1.60 \times 10^{-22}$  g**

5. An oxide of rhenium crystallizes with rhenium atoms at each of the corners and oxygen atoms in the middle of each edge of the cell. What is the formula of the oxide?

**Answer : ReO<sub>3</sub>**

- 6.a. A face-centered cubic cell contains X atoms at the corners of the cell and Y atoms at the faces. What is the empirical formula of the solid?

**Answer : XY<sub>3</sub>**

- b. The unit cell in a certain lattice consists of a cube formed by an anion, A, at each corner, an anion in the center, and a cation, X, at the center of each face. How many anions and cations are there in the unit cell?

**Answer: 3 cations and 2 anions**

7. Barium crystallizes in a body-centered cubic unit cell with a density of 3.62 g/cm<sup>3</sup>. Calculate the atomic radius of barium in picometers ?

Molar mass of barium is 137.3 g/mol .

# Chapter 12 Answers Natural Approach To Chemistry

**IM Harris**



## **Chapter 12 Answers Natural Approach To Chemistry:**

**The Natural Method** Eddy Nahmias, Thomas W. Polger, Wenqing Zhao, 2020-08-04 Prominent philosophers explore themes in the work of Owen Flanagan focusing on debates about the nature of mind the self and morality Owen Flanagan s work offers a model for how to be a naturalistic and scientifically informed philosopher who writes beautifully and deeply about topics as varied as consciousness and Buddhism moral psychology and dreaming identity and addiction literature and neuroscience In this volume leading philosophers Flanagan s friends colleagues and former students explore themes in his work focusing on debates over the nature of mind the self and morality Some contributors address Flanagan s work directly others are inspired by his work or methodology Their essays are variously penetrating and synoptic cautious and speculative

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**Nature** Sir Norman Lockyer,1908     Announcements for the Scholastic Year St. Procopius College,1961     **The New Statesman** ,1919     School Science and Mathematics ,1925     **The Chemical Trade Journal and Chemical Engineer** G Kelville Davis,1914     The Bookseller and the Stationery Trades' Journal ,1886 Official organ of the book trade of the United Kingdom     **Nature** ,1881     **Journal of Education and School World** ,1886

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