

12

STOICHIOMETRY

Practice Problems

In your notebook, solve the following problems.

SECTION 12.1 THE ARITHMETIC OF EQUATIONS

Use the 3-step problem-solving approach you learned in Chapter 1.

1. An apple pie needs 10 large apples, 2 cups rasp and blueberries, and 1 tablespoon of cinnamon. Write a balanced equation that fits this situation. How many apples are needed to make 20 pies?
2. Two moles of potassium chloride and three moles of oxygen are produced from the decomposition of two moles of potassium chlorate, KClO_3 . Write the balanced equation. How many moles of oxygen are produced from 12 moles of potassium chlorate?
3. Using the equation from problem 2, how many moles of oxygen are produced from 14 moles of potassium chlorate?
4. Two molecules of hydrogen react with one molecule of oxygen to produce two molecules of water. How many molecules of water are produced from 3.8×10^{22} molecules of oxygen? How many moles of water are produced from 22.5 moles of oxygen?

SECTION 12.2 CHEMICAL CALCULATIONS

1. Calculate the number of moles of hydrogen chloride produced from 10 moles of hydrogen.



2. Calculate the number of moles of chlorine needed to form 14 moles of iron(III) chloride.



3. Calculate the number of grams of nitrogen dioxide that are produced from 4 moles of nitric oxide.



4. Calculate the mass of oxygen produced from the decomposition of 15.0 g of potassium chlorate.



5. Calculate the mass of silver needed to react with chlorine to produce 94 g of silver chloride. (Write a balanced equation first.)

6. How many liters of carbon monoxide at STP are needed to react with 4.00 g of oxygen gas to produce carbon dioxide?



7. Calculate the number of liters of oxygen gas needed to produce 15.0 liters of nitrogen dioxide. Assume all gases are at the same conditions of temperature and pressure.



Chapter 12 Stoichiometry Reading Guide Answers

Martha B. Barrett



Chapter 12 Stoichiometry Reading Guide Answers:

The Practice of Chemistry Study Guide & Solutions Manual Pamela Mills, Amina El-Ashmawy, 2003-04-14 Designed to help students understand the material better and avoid common mistakes Also includes solutions and explanations to odd numbered exercises

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Table of Contents Chapter 12 Stoichiometry Reading Guide Answers

1. Understanding the eBook Chapter 12 Stoichiometry Reading Guide Answers
 - The Rise of Digital Reading Chapter 12 Stoichiometry Reading Guide Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 12 Stoichiometry Reading Guide Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chapter 12 Stoichiometry Reading Guide Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 12 Stoichiometry Reading Guide Answers
 - Personalized Recommendations
 - Chapter 12 Stoichiometry Reading Guide Answers User Reviews and Ratings
 - Chapter 12 Stoichiometry Reading Guide Answers and Bestseller Lists

5. Accessing Chapter 12 Stoichiometry Reading Guide Answers Free and Paid eBooks
 - Chapter 12 Stoichiometry Reading Guide Answers Public Domain eBooks
 - Chapter 12 Stoichiometry Reading Guide Answers eBook Subscription Services
 - Chapter 12 Stoichiometry Reading Guide Answers Budget-Friendly Options
6. Navigating Chapter 12 Stoichiometry Reading Guide Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Chapter 12 Stoichiometry Reading Guide Answers Compatibility with Devices
 - Chapter 12 Stoichiometry Reading Guide Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chapter 12 Stoichiometry Reading Guide Answers
 - Highlighting and Note-Taking Chapter 12 Stoichiometry Reading Guide Answers
 - Interactive Elements Chapter 12 Stoichiometry Reading Guide Answers
8. Staying Engaged with Chapter 12 Stoichiometry Reading Guide Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chapter 12 Stoichiometry Reading Guide Answers
9. Balancing eBooks and Physical Books Chapter 12 Stoichiometry Reading Guide Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter 12 Stoichiometry Reading Guide Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chapter 12 Stoichiometry Reading Guide Answers
 - Setting Reading Goals Chapter 12 Stoichiometry Reading Guide Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chapter 12 Stoichiometry Reading Guide Answers
 - Fact-Checking eBook Content of Chapter 12 Stoichiometry Reading Guide Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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