Question 13.13:

A gas in equilibrium has uniform density and pressure throughout its volume. This is strictly true only if there are no external influences. A gas column under gravity, for example, does not have uniform density (and pressure). As you might expect, its density decreases with height. The precise dependence is given by the so-called law of atmospheres $n_2 = n_1 \exp[-mg(h_2 - h_1)/kBT]$

Where n_2, n_1 refer to number density at heights $h2$ and $h1$ respectively. Use this relation to derive the equation for sedimentation equilibrium of a suspension in a liquid column:

$n_2 = n_1 \exp[-mg\, NA(\rho - P')(h_2 - h_1)/(\rho RT)]$

Where ρ is the density of the suspended particle, and ρ' that of surrounding medium. [N_A is Avogadro's number, and R the universal gas constant.]

[**Hint**: Use Archimedes principle to find the apparent weight of the suspended particle.]

Answer 13.13:

According to the law of atmospheres, we have:

$n_2 = n_1 \exp[-mg(h_2 - h_1)/kBT]$ ... (i)

Where,

n_1 is the number density at height h_1, and n_2 is the number density at height h_2.

mg is the weight of the particle suspended in the gas column.

Density of the medium = ρ'

Density of the suspended particle = ρ

Mass of one suspended particle = m'

Mass of the medium displaced = m

Volume of a suspended particle = V

According to Archimedes' principle for a particle suspended in a liquid column, the effective weight of the suspended particle is given as:

Weight of the medium displaced − Weight of the suspended particle

$= mg - m'g$

$= mg - V\rho' g = mg - \left(\frac{m}{\rho}\right)\rho' g$

$= mg\left(1 - \frac{\rho'}{\rho}\right)$ ...(ii)

Gas constant, $R = kBN$

$k_B = \frac{R}{N}$ .... (iii)

Substituting equation (ii) in place of mg in equation (i) and then using equation (iii), we get:

$n_2 = n_1 \exp[-mg(h_2 - h_1)/kBT]$

$= n_1 \exp\left[-mg\left(-\frac{\rho'}{\rho}\right)_{(h_2-h_1)}\frac{N}{RT}\right]$

$= n_1 \exp\left[-mg(\rho - \rho')_{(h_2-h_1)}\frac{N}{RT\rho}\right]$

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