

# Chapter 14: Acids and Bases



## Properties of Acids

- Sour taste
- React with some metals
- Turns blue litmus paper red
- React with bases



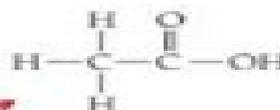
## Some Common Acids

HCl, hydrochloric acid

H<sub>2</sub>SO<sub>4</sub>, sulfuric acid

HNO<sub>3</sub>, nitric acid

HC<sub>2</sub>H<sub>3</sub>O<sub>2</sub>, acetic acid  
(common name)



What is the IUPAC name of this carboxylic acid?

## Properties of Bases

- Taste bitter
- Solutions feel slippery
- Turns red litmus paper blue
- React with acids

## Some Common Bases

NaOH, sodium hydroxide

KOH, potassium hydroxide

NaHCO<sub>3</sub>, sodium bicarbonate

## Arrhenius Acid-Base Theory

**Arrhenius Acid** - a hydrogen containing compound that ionizes to produce hydrogen ions (H<sup>+</sup>) when dissolved in water



Because **molecular acids are not made of ions**, they cannot dissociate.

They must be pulled apart, or **ionized**, by the water.

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**Christian G. Meyer**



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