balance nuclear reaction equations

aluminum-26 undergoes positron emission $^{26}_{13}Al \rightarrow ^{0}_{1}e + ^{26}_{12}Mg$

potassium-40 undergoes beta decay (electron) $^{40}_{19}K \rightarrow ^{0}_{-1}e + ^{40}_{20}Ca$

uranium-238 undergoes alpha decay $^{238}_{92}U \rightarrow ^{4}_{2}He + ^{234}_{90}Th$

electron configurations

Write the electron configuration for C. $1s^2 2s^2 2p^2$

Write the electron configuration for Cr. $1s^2 2s^2 2p^6 3s^2 3p^6 4s^1 3d^5$

Write the shorthand notation for Bi. $[Xe]\ 6s^2 4f^{14} 5d^{10} 6p^3$

NOMENCLATURE (naming compounds, writing formulas)

Write the formula for each of the following compounds:

a. Sulfur hexafluoride SF_6
b. Lithium nitride Li_3N
c. Chromium (III) carbonate $Cr_2(CO_3)_3$
d. Tin (II) chloride $SnCl_2$
e. Ammonium acetate $NH_4C_2H_3O_2$ (or CH_3COONH_4)
f. Mercury (I) chloride Hg_2Cl_2
g. Potassium bromate $KBrO_3$
h. Hydrosulfuric acid $H_2S_{(aq)}$
i. Chloric acid $HClO_3$
j. Sulfurous acid H_2SO_3

Name each of the following compounds.

a. $CuSO_4$ copper (II) sulfate
b. AlF_3 aluminum fluoride
c. HI hydrogen iodide
d. NO nitrogen monoxide
e. $H_2Se_{(aq)}$ hydroselenic acid
f. HNO_3 nitric acid
g. HNO_2 nitrous acid
h. $NaHSO_4$ sodium hydrogen sulfate

Half-life

Tritium (H-3) is a radioactive isotope of hydrogen with a half-life of 12.3 years. How long would it takes for a 40.0 g sample to decay down to 1.25 g? $40 \xrightarrow{1} 20 \xrightarrow{2} 10 \xrightarrow{3} 5 \xrightarrow{4} 2.5 \xrightarrow{5} 1.25$ $\boxed{61.5\ yr}$

Fe-61 has a half-life of 6.00 min. Of a 100.0 mg sample, how much will remain after 18.0 min?

3 half lives $100 \xrightarrow{1} 50 \xrightarrow{2} 25 \xrightarrow{3} \boxed{12.5\ mg}$

Average atomic mass

The element copper has naturally occurring isotopes with mass numbers of 63 and 65.

The relative abundance and atomic masses are 69.2% for a mass of 63amu and 30.8% for a mass of 65amu. Calculate the average atomic mass of copper.

$63 \times .692 = 43.596 \rightarrow 44$

$65 \times .308 = 20.02 \rightarrow 20.$

$+$

$63.62 \rightarrow 64$

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Sebastian Brünink

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