

Chapter 18- Radioactivity and Nuclear Reactions

I. Radioactivity

- A. The nucleus of an atom is composed of protons and neutrons which comprise most of the atom's mass.
- B. The strong force causes protons and neutrons to be attracted to each other in the nucleus.
1. The strong force is powerful only when neutrons and protons are closely packed together.
 2. The protons and neutrons in a large nucleus are held less tightly by the strong force than protons and neutrons in a small nucleus.
- C. Radioactivity—nuclear decay which happens when the strong force is not large enough to hold the nucleus together; the nucleus gives off matter and energy.
1. Isotopes are atoms of the same element with varying numbers of neutrons.
 2. A nucleus with too many or too few neutrons compared to protons is radioactive.
 3. Atomic number—number of protons in an atom
 4. Mass number—number of protons and neutrons in a nucleus
- D. Uranium radioactivity was discovered in 1896 by Henri Becquerel; Marie and Pierre Curie discovered the radioactive elements polonium and radium in 1898.

II. Nuclear Decay

- A. Nuclear radiation—particles and energy are released from a decaying nucleus.
- B. Alpha particle—two protons and two neutrons with an electric charge of +2
1. Alpha particles leave charged ions in their path when they travel through matter.
 2. Alpha particles are the least penetrating form of nuclear radiation.
 3. Alpha particles can cause serious biological damage.
- C. Beta particle—neutron decays into a proton and releases an electron at high speed; more penetrating than alpha particles
- D. Gamma ray—penetrating electromagnetic waves that carry energy, but have no mass or charge
- E. Transmutation—process of one element's changing to another through nuclear decay
- F. The half-life of a radioactive isotope is the length of time it takes half the nucleus to decay; half-lives vary from fractions of a second to billions of years.
- G. Carbon dating can be used to date once-living materials while uranium dating can be used to date rocks.

III. Detecting Radioactivity

- A. Radiation detectors are instruments used to identify ions formed when radiation passes through matter.
1. A cloud chamber detects alpha or beta particles by means of a trail of condensed vapor.
 2. A bubble chamber detects radioactive particles by means of a bubble trail in a superheated liquid.
 3. Electroscopes can measure charged particles in the air.
- B. A Geiger counter measures radioactivity by producing an electric current when radiation is present.
- C. Background radiation comes from radon gas produced in the Earth's crust, from cosmic rays, and from radioactive isotopes in the body.

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