Chemical Equilibrium

Teacher Notes and Answers

CONCEPT 10

REACTIONS 1

REVIEW QUESTIONS

1. a. $K_c = \frac{[C]^2[D]^3}{[A][B]^2}$

   b. $K_c = \frac{[NH_3]^2}{[N_2][H_2]^3}$

   (i) $K_c = \frac{[CO_2][H_2O]}{[CO][H_2O]} = 1.0 \times 10^{-4}$

   (ii) $K_c = \frac{[CO_2][H_2O]}{[CO][H_2O]} = 1.0 \times 10^{-4}$

   (iii) $K_c = \frac{[CO_2]}{[CO]} = 1.0 \times 10^{-4}$

   (iv) $K_c = \frac{[CO_2]^2}{[CO]^2[H_2O]} = 1.0 \times 10^{-4}$

   c. (i) Increase

   (ii) Decrease

   (iii) Decrease

   d. (i) The ratio of the concentrations of the products to the reactants remains the same as the ratio in the new flask.

   (ii) It is a coincidence that the products in both the initial flask and the new flask have the same numbers of moles.

2. a. 2.0

   b. 0.0 Molar

REACTIONS 2

REVIEW QUESTIONS

1. a. K_c

   b. K_c

   (i) K_c

   (ii) K_c

   c. (i) Decreases

   (ii) Increases (Becomes more)

   (iii) Stays the same

   (iv) Decreases

2. a. K_c

   b. Those results are only valid for a specific temperature. Changing the temperature changes the equilibrium constant. Which way it changes depends on the reaction. The equilibrium constant is a function of temperature.

   c. The equilibrium constant is independent of the initial concentrations of the reactants and products. It depends only on the temperature.

   d. Changing a hydrogen bond in the reactants or products changes the equilibrium constant. Hydrogen bonds are very strong and have a significant effect on the equilibrium constant.

   e. Adding a catalyst does not change the equilibrium constant.

   f. Adding a catalyst does not change the equilibrium position.

   g. Adding a catalyst does not change the equilibrium constant. When a catalyst is added, the equilibrium position does not change. The equilibrium constant is independent of the initial concentrations of the reactants and products. It depends only on the temperature.

   h. Changing the concentration of the reactants or products changes the equilibrium position. It does not change the equilibrium constant. The equilibrium constant is independent of the initial concentrations of the reactants and products. It depends only on the temperature.

REACTIONS 3

REVIEW QUESTIONS

1. a.
2. a.
3. a.
4. a.
5. a.
6. a. (i) Shift to the left (ii) No shift (iii) Shift to the left (iv) Shift to the left

   (v) Shift to the left (vi) No shift (vii) Shift to the left

   b. (i)

   (ii)

   c. If the reaction is being carried out in a closed flask, the equilibrium position does not change.

7. a. $K_{sp} = \frac{[Ag^+][Br^-]}{[AgBr]}$

   b. (i) Shift to the left (ii) No shift

   (iii) $K_{sp} = \frac{[Ag^+][Br^-]}{[AgBr]}$

   (iv) $K_{sp} = 5.0 \times 10^{-13}$

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Hussin A.Rothana

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