

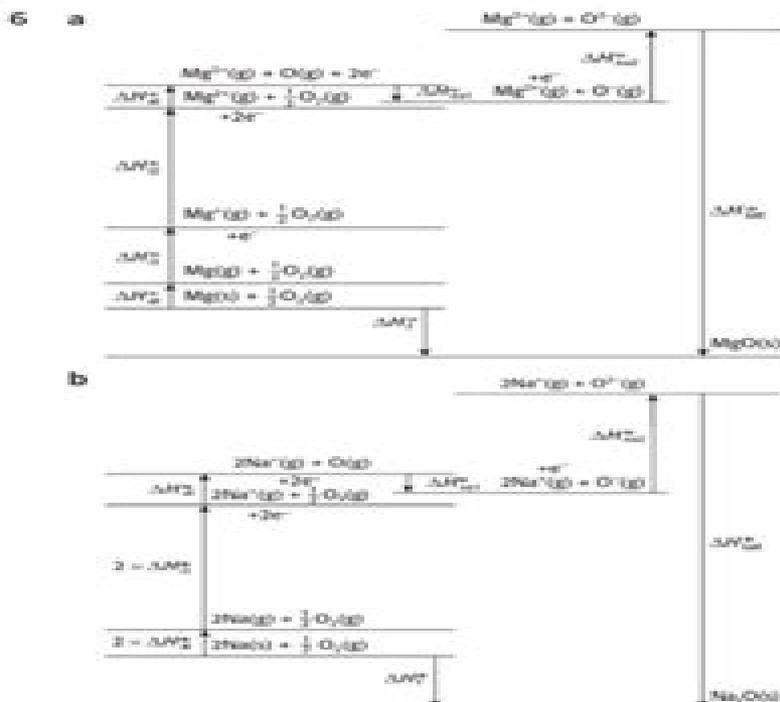
Answers to SAQs

Chapter 19

- 1 a standard temperature = 298 K, standard pressure = 1×10^5 Pa / 100 kPa
- b i $\text{Mg}^{2+}(\text{g}) + \text{O}^{2-}(\text{g}) \rightarrow \text{MgO}(\text{s})$
 ii $\text{K}^+(\text{g}) + \text{Br}^-(\text{g}) \rightarrow \text{KBr}(\text{s})$
 iii $2\text{Na}^+(\text{g}) + \text{S}^{2-}(\text{g}) \rightarrow \text{Na}_2\text{S}(\text{s})$
- 2 a The bond energy for chlorine is the enthalpy change $\text{Cl}_2(\text{g}) \rightarrow 2\text{Cl}(\text{g})$. The enthalpy change of atomisation is $\frac{1}{2}\text{Cl}_2(\text{g}) \rightarrow \text{Cl}(\text{g})$. So the enthalpy change of atomisation is $+244/2 = +122 \text{ kJ mol}^{-1}$.
- b i $\frac{1}{2}\text{O}_2(\text{g}) \rightarrow \text{O}(\text{g})$
 ii $\text{Ba}(\text{s}) \rightarrow \text{Ba}(\text{g})$
 iii $\frac{1}{2}\text{Br}_2(\text{l}) \rightarrow \text{Br}(\text{g})$
- c 0 kJ mol^{-1}
 Because helium exists naturally as single gaseous atoms, no change is involved in the process $\text{He}(\text{g}) \rightarrow \text{He}(\text{g})$.
- 3 a There must be an input of energy to overcome the repulsive forces between the (negative) electron and the negative ion.
- b $+440 \text{ kJ mol}^{-1}$
- c i $\text{I}(\text{g}) + \text{e}^- \rightarrow \text{I}^-(\text{g})$
 ii $\text{S}^-(\text{g}) + \text{e}^- \rightarrow \text{S}^{2-}(\text{g})$
- 4 a i $\text{Cs}(\text{g}) \rightarrow \text{Cs}^+(\text{g}) + \text{e}^-$
 ii $\text{Al}^{3+}(\text{g}) \rightarrow \text{Al}^{2+}(\text{g}) + \text{e}^-$
 iii $\text{Ca}(\text{s}) + \frac{1}{2}\text{O}_2(\text{g}) \rightarrow \text{CaO}(\text{s})$
 iv $\text{Fe}(\text{s}) + 1\frac{1}{2}\text{Cl}_2(\text{g}) \rightarrow \text{FeCl}_3(\text{s})$
- b $\Delta H_{\text{form}}^{\ominus} = \Delta H_{\text{at}}^{\ominus} - [\Delta H_{\text{ion}}^{\ominus}(\text{Na}) + \Delta H_{\text{ion}}^{\ominus}(\text{Cl}) + \Delta H_{\text{at}}^{\ominus}(\frac{1}{2}\text{Cl}_2(\text{g})) + \Delta H_{\text{form}}^{\ominus}(\text{Cl})]$
 $\Delta H_{\text{form}}^{\ominus} = (-411) - \{(+107) + (+496) + (+122) + (-348)\}$
 $= +377 \text{ kJ mol}^{-1}$
 $\Delta H_{\text{form}}^{\ominus} = (-411) - (+377) = -788 \text{ kJ mol}^{-1}$



- b i enthalpy change of atomisation of iodine
 ii first electron affinity of nitrogen
 iii enthalpy change of formation of strontium chloride
 iv lattice energy of cadmium chloride



Chapter 19 Answer Key Chemistry

Lauren Gardner



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