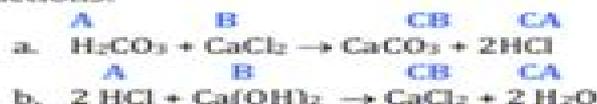


Chapter 19 Study Guide: Acids, Bases & Titration – Answers

1. Differentiate between Arrhenius, Brønsted-Lowry & Lewis acids **and** bases.

	ACIDS	BASES
Arrhenius	Form hydronium ions (H_3O^+) in aqueous sol'n	Form hydroxide ions (OH^-) in aqueous sol'n
Brønsted-Lowry	Proton (H^+) donors	Proton (H^+) acceptors
Lewis	Electron pair acceptors	Electron pair donors

2. Label the acid (A), base (B), conjugate acid (CA) and conjugate base (CB) in each of the following reactions:



3. Write the names or formulas for the following acids and bases:

- | | |
|--|---|
| a. H_2S – hydrosulfuric acid | e. Nitrous acid – HNO_2 |
| b. $\text{HC}_2\text{H}_3\text{O}_2$ – acetic acid | f. Calcium hydroxide – $\text{Ca}(\text{OH})_2$ |
| c. NaOH – sodium hydroxide | g. hydrochloric acid – HCl |
| d. HNO_3 – nitric acid | |

Use the following formulas for the remaining questions: (these will not be given to you on the test!)

$$K_w = [\text{H}_3\text{O}^+][\text{OH}^-] = 1.0 \times 10^{-14}$$

$$\text{pH} = -\log[\text{H}_3\text{O}^+]$$

$$\text{pOH} = -\log[\text{OH}^-]$$

$$\text{pH} + \text{pOH} = 14$$

$$M_A V_A = n_A = M_B V_B = n_B$$

4. What is the pH of 1.34×10^{-4} M solution of hydrochloric acid?

$$\text{pH} = -\log [1.34 \times 10^{-4} \text{M}] = 3.873$$

5. Determine the pH and pOH of a 1.85×10^{-6} M NaOH solution.

$$\text{pOH} = -\log [1.85 \times 10^{-6} \text{M}] = 5.733$$

$$\text{pH} = 14 - 5.73 = 8.267$$

6. What is the $[\text{OH}^-]$ of an aqueous solution that has a pH of 5.0?

$$\text{pOH} = 9 \quad [\text{OH}^-] = 1 \times 10^{-9} \text{ M}$$

7. What is the pH and the pOH of a solution with a volume of 2000. mL containing 35 grams of NaOH? Is this solution acidic or basic?

$$\frac{35 \text{g NaOH}}{2.000 \text{L}} \times \frac{1 \text{mol NaOH}}{40.00 \text{g NaOH}} = 0.44 \text{ M NaOH}$$

$$\text{pOH} = -\log[0.44 \text{M}] = 0.36 \quad \text{pH} = 14 - 0.36 = 13.64 \quad \text{basic}$$

Chapter 19 Study Guide Acids Bases Answers

CL Gary



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Table of Contents Chapter 19 Study Guide Acids Bases Answers

1. Understanding the eBook Chapter 19 Study Guide Acids Bases Answers
 - The Rise of Digital Reading Chapter 19 Study Guide Acids Bases Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 19 Study Guide Acids Bases Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chapter 19 Study Guide Acids Bases Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 19 Study Guide Acids Bases Answers
 - Personalized Recommendations
 - Chapter 19 Study Guide Acids Bases Answers User Reviews and Ratings
 - Chapter 19 Study Guide Acids Bases Answers and Bestseller Lists
5. Accessing Chapter 19 Study Guide Acids Bases Answers Free and Paid eBooks
 - Chapter 19 Study Guide Acids Bases Answers Public Domain eBooks
 - Chapter 19 Study Guide Acids Bases Answers eBook Subscription Services
 - Chapter 19 Study Guide Acids Bases Answers Budget-Friendly Options
6. Navigating Chapter 19 Study Guide Acids Bases Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Chapter 19 Study Guide Acids Bases Answers Compatibility with Devices
 - Chapter 19 Study Guide Acids Bases Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chapter 19 Study Guide Acids Bases Answers
 - Highlighting and Note-Taking Chapter 19 Study Guide Acids Bases Answers
 - Interactive Elements Chapter 19 Study Guide Acids Bases Answers
8. Staying Engaged with Chapter 19 Study Guide Acids Bases Answers

- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chapter 19 Study Guide Acids Bases Answers
9. Balancing eBooks and Physical Books Chapter 19 Study Guide Acids Bases Answers
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter 19 Study Guide Acids Bases Answers
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chapter 19 Study Guide Acids Bases Answers
- Setting Reading Goals Chapter 19 Study Guide Acids Bases Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chapter 19 Study Guide Acids Bases Answers
- Fact-Checking eBook Content of Chapter 19 Study Guide Acids Bases Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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