

ACIDS, BASES AND SALTS

Definition of an acid An acid is a compound which when dissolved in water produces hydrogen ions as the only positively charged ions	What causes the acidic properties of acids? The hydrogen ions (H ⁺) cause the acidic properties and these are formed in the presence of water.
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Another term to refer to an acid An acid is called a <u>proton donor</u>	Why an acid is also called a proton donor? It's because an acid provides protons or hydrogen ions (H ⁺) to other substances during the reaction.
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Substance to which an acid provides protons BASE	A base is then called PROTON ACCEPTOR	Why the base is called a proton acceptor? Because the base accepts hydrogen ions from acids	Equation for the reaction between acid and base $H^+_{(aq)} + OH^-_{(aq)} \longrightarrow H_2O_{(l)}$
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Common laboratory acids	These three common laboratory acids are also called	Why these acids are also called mineral acids?
Hydrochloric acid Sulphuric acid Nitric acid	HCl H ₂ SO ₄ HNO ₃ MINERAL ACIDS	They are derived from mineral salts i.e chlorides for HCl, sulphates for H ₂ SO ₄ and nitrates for HNO ₃

Other mineral acids known	Mineral salts from which the acid is derived	Organic acids known	Naturally occurring acids
Sulphurous acid Carbonic acid Phosphoric acid Nitrous acid	H ₂ SO ₃ H ₂ CO ₃ H ₃ PO ₄ HNO ₂ Derived from sulphites Derived from carbonates Derived from phosphates Derived from nitrites	Ethanoic acid (CH ₃ COOH) Methanoic acid (HCOOH)	CITRIC ACID from lemons TARTARIC ACID from grapes ACETIC ACID from vinegar LACTIC ACID from sour milk Hydrochloric acid from digestive juices

Whenever an acid is dissolved in water, it produces HYDROGEN IONS	Term given to the number of hydrogen ions produced by one molecule of an acid BASICITY OF AN ACID	Definition of basicity of an acid BASICITY of an acid is the <u>number of hydrogen ions</u> produced by <u>one molecule</u> of an acid <u>in aqueous solution</u> .
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Chapter 1acids And Bases Study Guide

T. W. Graham Solomons



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