

## CHEMISTRY SEMESTER 1 REVIEW QUESTIONS KEY

These questions are a small selection for each unit this semester. Refer to the exam scoring-guide and your self-reflection guide to determine which questions you should focus your studies on. Answers to select questions can be found at the end of this document.

### Unit 1: Gas Laws

1. A gas is heated in a closed, rigid vessel. What happens to the molecules inside?  
The molecules start to move faster. They hit the inside of the container with greater force, which increases the pressure.
2. A gas is compressed (volume decreased) at a constant temperature. What happens to the pressure inside the container? Explain your reasoning.  
Since the volume is decreased, the molecules hit the sides of the container with increased frequency. This raises the pressure.
3. When inflating a car tire, the pressure doubles from 200 to 400 kPa. Why would the pressure double when the volume is increasing?  
The volume may be increasing, but it is because more molecules are being added to the system. More molecules means more collisions, which increases the pressure.
4. What is the relation between temperature and pressure when volume is held constant?  
With a constant volume, if temperature increases pressure increases as well. Think about heating an aerosol can (hair spray) – the volume is rigid and fixed. If you heat the can, the internal molecules move faster, hitting the sides harder and more frequently. This will increase the pressure so much that the can may explode.

### Unit 2: Atomic structure and theory, periodic table

5. How many protons, neutron, and electrons are in a neutral atom of Ag-108?  
Protons = 47  
Neutrons = 61  
Electrons = 47

6. Fill in the table below.

Name	Sodium	Nitrogen
Symbol	$^{23}_{11}\text{Na}$	$^{14}_7\text{N}$
Atomic number	11	7
Mass number	23	15
Protons	11	7
Neutrons	12	8
Electrons in a neutral atom	11	7

7. Analyze the following two isotopes. What is the same between both isotopes? What is different?  $^{15}_7\text{N}$  and  $^{14}_7\text{N}$   
Both isotopes have the same number of protons (7) yet a different number of neutrons (8 and 7, respectively). The different number of neutrons yields a different mass number for the isotopes.
8. Chlorine has 2 isotopes, Cl-35 and Cl-37. Cl-35 has an abundance of 75.77% while Cl-37 has an abundance of 24.23%. Will the atomic mass of chlorine be closer to 35 or 37? Explain your reasoning?  
Atomic mass will be closer to Cl-35 because 'atomic mass' refers to the average mass of all isotopes of an element. If most chlorine isotopes have a mass of 35 while only a few have a mass of 37, the average mass will be closer to 35.
9. Draw the Bohr model for sulfur. Label protons, neutrons, electrons, nucleus, and energy levels.



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**S Ashworth**



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