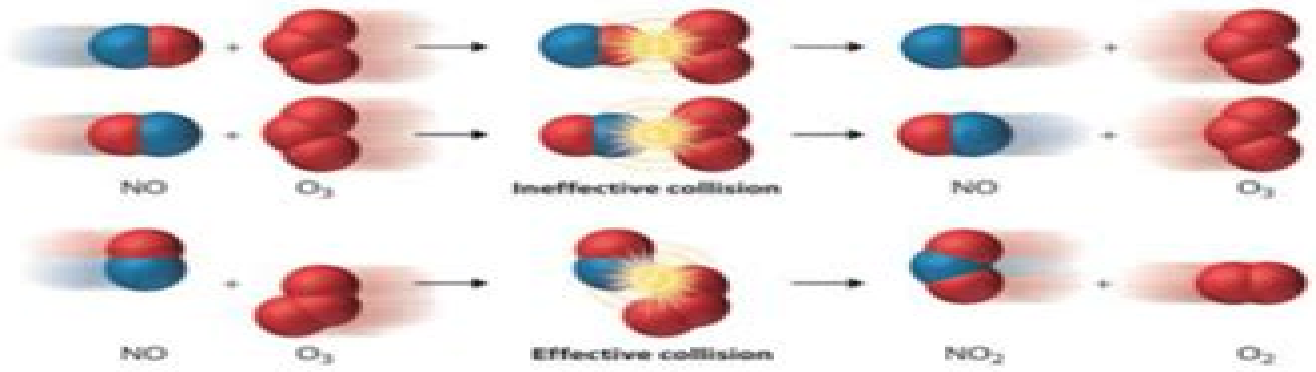
Topic	Collision theory	Level	GCSE
Outcomes	1. To use the collision theory to explain why changes to a reaction can increase the rate of reaction (temperature, concentration, surface area and catalyst) 2. To understand why not all collisions are successful		

The collision theory and rates of reaction



When two particles collide (hit each other) sometimes a chemical reaction happens. Not all collisions result in a chemical reaction. To increase the number of successful collisions we can change the conditions of a reaction (see page 2).

1. Do all collisions between particles cause a reaction to happen?
2. What do we call the minimum energy the particles must have for a reaction to happen?
3. What do we call a collision between two particles that causes a chemical reaction?
4. Look at the picture above, how do you know the first reaction is not a successful collision?

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