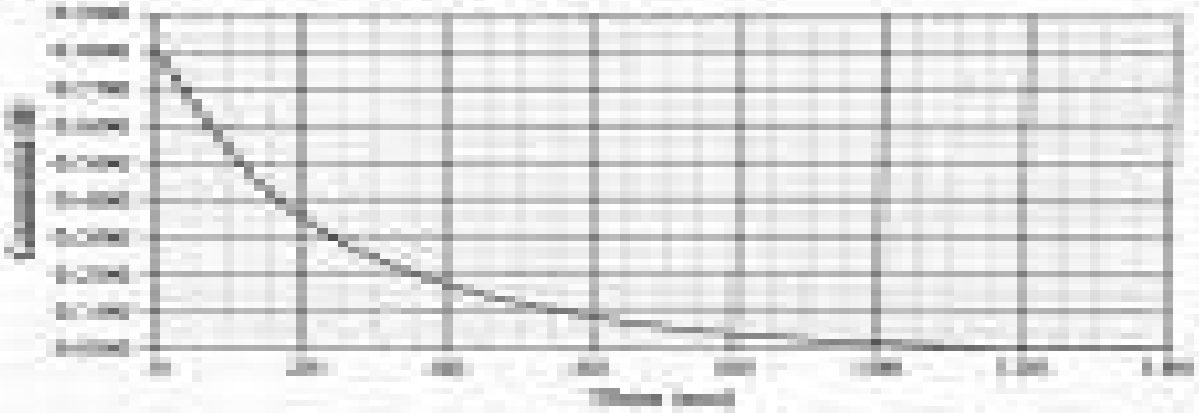
Rate of Reaction

How is the speed of a reaction measured?

Why?

Chemical reactions occur at different speeds. Some are almost instantaneous. Other require patience. For example, rust can form in days or years. Rates are used to predict if reactions happen. Depending on the conditions, it could be nearly impossible to change the speed of a reaction, so most time spending on understanding of how the rate of reaction is measured during a reaction.

Model 1 — Concentration versus Time Graph



1. The graph in Model 1 illustrates how the concentration of a species in a chemical reaction changes over time.
   a. Where unit is the concentration measured in?
   The concentration is measured in molarity.
   b. Where unit is the time measured in?
   The time is measured in seconds.
2. Consider the data in Model 1:
   a. What was the concentration of the species when the chemical reaction was initiated?
   The initial concentration of the chemical species in the reaction was 0.80 M.
   b. Did the concentration of the species increase or decrease over time?
   The concentration of the chemical species decreased over time.
   c. Was the species a reactant or product in the reaction? Briefly explain reasoning.
   The chemical species was a reactant because it was used up during the reaction. Its concentration went down.

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