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distributed by the College Board Shortchanged Annie Abrams, 2023-04-25 Shortchanged is a brilliant book The Washington Post Author and high school English teacher Annie Abrams reveals how the College Board's emphasis on standardized testing has led the AP program astray Every year millions of students take Advanced Placement AP exams hoping to score enough points to earn college credit and save on their tuition bill But are they getting a real college education The College Board says that AP classes and exams make the AP program more accessible and represent a step forward for educational justice But the program's commitment to standardized testing no longer reflects its original promise of delivering meaningful college level curriculum to high school students In Shortchanged education scholar Annie Abrams uncovers the political and pedagogical traditions that led to the program's development in the 1950s In revealing the founders' intentions of aligning liberal arts education across high schools and colleges in ways they believed would protect democracy Abrams questions the collateral damage caused by moving away from this vision The AP program is the College Board's greatest source of revenue yet its financial success belies the founding principles it has abandoned Instead of arguing for a wholesale restoration of the program Shortchanged considers the nation's contemporary needs Abrams advocates for broader access to the liberal arts through robust public funding of secondary and higher education and a dismantling of the standardized testing regime Shortchanged illuminates a better way to offer a quality liberal arts education to high school students while preparing them for college

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web 1 molarity  $M$  mass molar mass  $\times 0.4500 \text{ l} = 0.825 \text{ g}$  141 9579  $\text{g mol}^{-1} \times 0.0129 \text{ m} = 2 \text{ molality}$  0.825  $\text{g}$  141 9579  $\text{g mol}^{-1}$  00581158  $\text{mol}^{-1}$  00581158  $\text{mol}^{-1}$  0.4500  $\text{kg}$  0.0129  $\text{m} = 3 \text{ mole fraction}$   $n_2$   $\text{H}_2\text{O}$  4.0825  $\text{g}$  141 9579  $\text{g mol}^{-1}$  00581158  $\text{mol}^{-1}$   $h_2o$  450.0  $\text{g}$  18.015  $\text{g mol}^{-1}$  24.97918401  $\text{mol}$

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web molality differs from molarity only in the denominator while molarity is based on the liters of solution molality is based on the kilograms of solvent concentrations expressed in molality are used when studying properties of solutions related to vapor pressure and temperature changes

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web what mass of  $\text{CoCl}_2$  in grams is needed for the solution express the answer using 3 significant figures

**8 3 concentrations of solutions problems chemistry libretxts** - Dec 13 2022

web the hardness of water hardness count is usually expressed in parts per million by mass of  $\text{CaCO}_3$  which is equivalent to milligrams of  $\text{CaCO}_3$  per liter of water what is the molar concentration of  $\text{Ca}^{2+}$  ions in a water sample with a hardness count of 175  $\text{mg CaCO}_3 \text{ l}^{-1}$

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web answer the mass percentage of an element of a solution is defined as the mass of the solute in grammes present in 100 g of the solution it is written like this  $\frac{\text{mass of a component}}{\text{mass of component in solution total mass of solution}} \times 100$

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web aug 13 2017 molarity practice problems this chemistry video tutorial explains how to solve common molarity problems it discusses how to calculate the concentration of a solution given the mass in grams