

CHEMISTRY



Periodic Table of Elements

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- Elemental forms:**
 - Monatomic:** Noble gases; helium, neon, argon, krypton, xenon, and radon; nonreactive
 - Diatoms:** Gases: fluorine, chlorine, hydrogen, oxygen, and nitrogen; Liquid: bromine; Solid: iodine
 - Molecular solids:** Sulfur
 - Covalent solids:** Carbon (diamond, graphite)
 - Metals:** Most common type of element; high coordination numbers; very dense
- Types of compounds:** Large variety due to ionic, covalent, and metallic bonding options
 - Ionic:** Electron exchange
 - Covalent:** Shared electrons
 - Polar covalent:** Unequal sharing
 - Ion complex:** Central ion + Solvent
 - Transition metal compounds:**
 - Coordination compound:** Non-bonding ligands
 - Organometallic:** Metal-C bond
 - Complex minerals:** Metal oxides and silicates
- Periodicity:** The chemical properties of an element depend on valence electrons
 - Family of elements:** Column in periodic table
 - The addition of d-orbitals produces an expanded octet

- Ionization energies:** Noble gases are the most stable; alkali metals are the easiest to ionize; electron screening lessens Z-effect on orbital energy
- Atomic and ionic radii:** Not real properties; derive from structural data; divide bond into radii using a consistent scheme
 - Ionic radii:** Depends on charge, counter ion, and coordination (crystal type)
 - Atomic radii:** Based on diatomic gas data or metallic solid structural data

Development of the Periodic Table

- Missing elements were predicted using periodicity
- Column 1:** Alkali metals
- Column 2:** Alkaline earth metals
- First transition series:** Titanium-Copper
- Column 17:** Halogens: fluorine, chlorine, bromine, iodine, and astatine
- Column 18:** Rare gases: helium, neon, argon, krypton, xenon, and radon
- Actinides:** Z = 90 or above; radioactive
- Lanthanides:** f-subshell elements; starts with Z = 58



Practical Application
 The colors of fireworks depend on the elements in the explosive



Atomic Structure

- Atomic number (Z):** The number of protons in the nucleus; for a neutral atom, Z = number of electrons; for an anion of charge -q, Z = q electrons; for a cation of charge +q, Z = q electrons
- Mass number (A):** A = Z + N, where N = the number of neutrons in the nucleus
- Isotopes:** Atoms with the same Z, different A; an element can have one or more isotopes; any sample contains a number of isotopes; in practical work, use the average mass number or atomic weight (weighted average of natural isotopes for a given element)



Atomic Quantum Numbers & Orbitals

- n:** principle
- l:** angular momentum (orbital shape)
- m_l:** magnetic (orbital direction)
- m_s:** electron spin
- For each n, the possible l values are 0, 1, ..., n - 1
- For each l, the possible m_l values are -l, 0, ..., +l
- m_s has two possible values: +1/2 and -1/2 (spin up, spin down)
- Each value of n denotes a "shell" in the atomic structure; each l denotes a subshell
 - s-orbital: l = 0, 1 type
 - p-orbital: l = 1, 3 types
 - d-orbital: l = 2, 5 types
 - f-orbital: l = 3, 7 types

Many-Electron Atoms

Orbitals and quantum numbers derived for the hydrogen atom are used to describe all many-electron atoms and ions; a more rigorous treatment determines exact energy levels and orbital properties; the main result is that subshells have different energies; the filling of levels is guided by the Aufbau principle:



n	l	m _l	m _s	element	orbital
1	0	0	+1/2	H	1s-orbital
1	0	0	-1/2	He	spherical
2	0	0	+1/2	Li	2s-orbital
2	0	0	-1/2	Be	spherical
2	1	-1	+1/2	B	2p-orbitals
2	1	0	+1/2	C	
2	1	0	-1/2	N	hemisidal
2	1	-1	+1/2	O	p, p, p
2	1	-1	-1/2	F	
2	1	+1	+1/2	Ne	

- Electron configuration:** Atomic orbital occupancy for an atom
- Pauli exclusion principle:** Each electron has a unique set of quantum numbers; an orbital may hold up to two electrons; one with spin up, m_s = +1/2, and one with spin down, m_s = -1/2
- Hund's rule:** Electrons fill p, d, and f subshells in a manner to maximize the number of unpaired electrons (half-filled orbitals); spin is maximized
- Ionization potential (IP):** Energy required to remove electron from an atom or ion
 - First IP:** Removal of first valence electron
- Electronegativity:** Tendency of an atom to attract electrons in a chemical bond; 0-4 scale; 0 for rare gas

College Chemistry Study Guide

JE Gale



College Chemistry Study Guide:

Unveiling the Power of Verbal Art: An Mental Sojourn through **College Chemistry Study Guide**

In a global inundated with screens and the cacophony of immediate interaction, the profound power and emotional resonance of verbal art frequently disappear in to obscurity, eclipsed by the regular barrage of sound and distractions. Yet, nestled within the musical pages of **College Chemistry Study Guide**, a fascinating work of fictional brilliance that impulses with natural emotions, lies an wonderful trip waiting to be embarked upon. Composed by way of a virtuoso wordsmith, this mesmerizing opus instructions visitors on an emotional odyssey, gently revealing the latent possible and profound impact embedded within the intricate web of language. Within the heart-wrenching expanse of this evocative analysis, we will embark upon an introspective exploration of the book is central themes, dissect their interesting publishing fashion, and immerse ourselves in the indelible impression it leaves upon the depths of readers souls.

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